

Daily Tutorial Sheet-14
Level - 3

153.(B) As the electronegativity of central atom E increases, the bond angle also increases.

154.(B) Bond order

He_2^+ 0.5

O_2^- 1.5

NO 2.5

C_2^{2-} 3

155.(D) BCl_3 has trigonal planar structure.

***156.(ABC)** $\text{NO} \rightarrow 11$ valence electrons $\text{NO}_2 \rightarrow 17$ valence electrons
 $\text{ClO}_2 \rightarrow 19$ valence electrons $\text{N}_2\text{O}_4 \rightarrow 34$ valence electrons

157(ACD) $\text{H}_2\overset{1}{\text{C}} = \overset{2}{\text{C}} = \overset{3}{\text{CH}_2}$

C-1 and C-3 lie in perpendicular planes. So, the whole molecule is non-planar.

158.(B) H_2O is $\text{H}-\overset{\cdot\cdot}{\underset{\cdot\cdot}{\text{O}}}-\text{H}$